

### **chapter 6 chemical bonds pdf**

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### **Chapter 6 Practice Test: Chemical Bonding - PDF documents**

Types of Chemical Bonding A chemical bond is a mutual electrical attraction between the nuclei and valence electrons of different atoms that binds the atoms together. When atoms form a chemical bond, their valence electrons are redistributed to make the atoms more stable. The way the electrons are redistributed determines the type of bond.

### **CHAPTER 6 Chemical Bonding - St. Charles Parish**

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### **Chapter 6 Chemical Bonding Quiz - PDF documents**

CHAPTER 6 REVIEW Chemical Bonding SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A covalent bond consists of (a) a shared electron.

### **6 Chemical Bonding - Effingham County Schools / Overview**

Test Bank - Chapter 6 . 6. If more reactants are used in a chemical reaction, more products will be ... b) More energy is released when the bonds in the products are formed than ...

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Chemical bonding that results from the electrical attraction between large numbers of cations and anions is called ionic bonding. In purely ionic bonding, atoms completely give up electrons to other atoms, as illustrated in Figure 6-1 on page 162. In contrast to atoms joined by ionic bonding, atoms joined by covalent bonding share elec-trons.

### **CHAPTER 6 Chemical Bonding**

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Chapter 6: Chemical Bonding Learning Objectives Describe the formation of ions by electron loss/gain to

obtain the electronic configuration of a noble gas. Describe the formation of ionic bonds between metals and non-metals E.g. NaCl State that ionic materials contain a giant lattice in which the ions are held by electrostatic attraction.

### **Chapter 6: Chemical Bonding - World of Chemistry**

6 A. Double Bonds 1. A covalent bond produced by the sharing of two pairs of electrons between two atoms 2. Higher bond energy and shorter bond length than single bonds B. Triple Bonds 1. A covalent bond produced by the sharing of three pairs of electrons between two atoms 2.

### **Chapter 6 Notes - srvhs.org**

6 single bonds (6x2) = 12 18 lone pairs (18x2) = 36 Total = 48 Dative Covalent Bond / Coordinate Covalent Bond A covalent bond in which one of the atoms donates both electrons. Examples: - NH<sub>4</sub><sup>+</sup>, NH<sub>3</sub>AlCl<sub>3</sub>, NH<sub>3</sub>BF<sub>3</sub> Hybridization Hybridization " mixing of two or more atomic orbitals to form a new set of hybrid orbitals. 1.

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